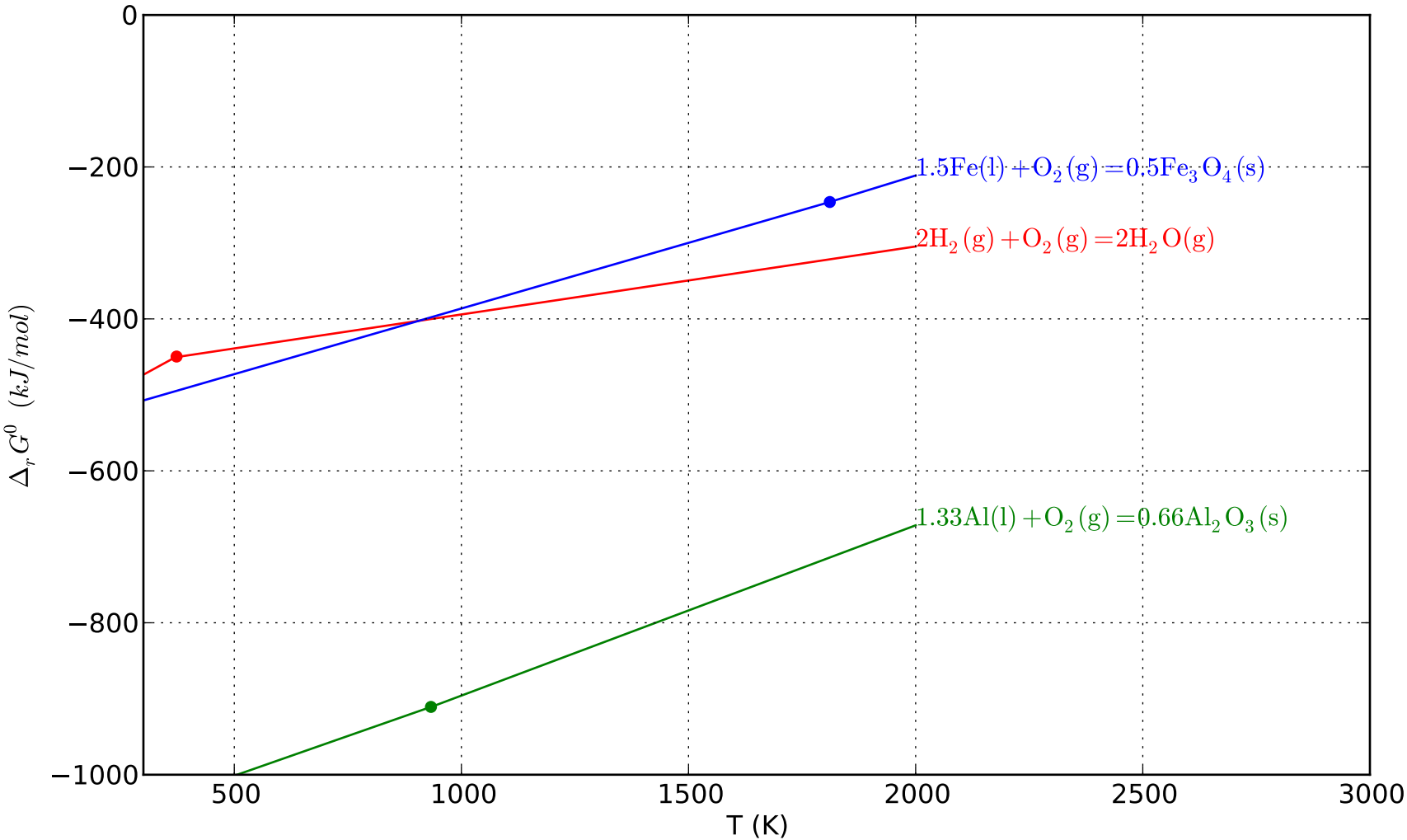
0
−200
−400
−600
−800
−1000
500
1000
1500
2000
2500
3000
T (K)
$\Delta_r G^0$ (kJ/mol)
$1.5Fe(l) + O_2(g) = 0.5Fe_3O_4(s)$
$2H_2(g) + O_2(g) = 2H_2O(g)$
$1.33Al(l) + O_2(g) = 0.66Al_2O_3(s)$